PERIODIC TABLE TRENDS WORKSHEET #2

Circle the correct element.

Li	Si	S	metal
N	P	As	smallest ionization energy
K	Ca	Sc	largest atomic mass
S	Cl	Ar	member of the halogen family
Al	Si	P	greatest electronegativity
Ga	Al	Si	largest atomic radius
V	Nb	Ta	largest atomic number
Te	I	Xe	member of noble gases
Si	Ge	Sn	4 energy levels
Li	Be	B	member of alkali metals
As	Se	Br	6 valence electrons
H	Li	Na	nonmetal
Hg	Tl	Pb	member of transition metals
Na	Mg	Al	electron config. ending in s^2p^1
Pb	Bi	Po	metalloid
B	C	N	gas at room temperature
Ca	Sc	Ti	electron config. ending in s^2d^2

Answers on your notebook pape:

1) Rank by increasing atomic radius:
carbon, aluminum, oxygen, potassium.

2) Rank by increasing electronegativity:
sulfur, oxygen, neon, aluminum.

3) Why does fluorine have a higher ionization energy than iodine?

4) Why do elements in the same family generally have similar properties?

5) Rank the sets of atoms from smallest to largest atomic radius.
a. Li, C, F b. Li, Na, K
c. Ge, P, O d. C, N, Al

6) Rank each set of atoms from lowest to highest ionization energy.
a. Mg, Si, S b. Mg, Ca, Ba
c. F, Cl, Br d. Ba, Cu, Ne e. Si, P, He

7) Rank each set of atoms from highest to lowest electronegativity.
a. Li, C, N b. C, O, Ne c. Si, P, O
d. K, Mg, P e. S, F, He

8) Brainstorm a mnemonic to help you remember which way the three trends (radius, ionization energy, electronegativity) increase on the PT (up/down/left/right)

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